

NAMES, FORMULAS AND CHARGES OF COMMON IONS

Positive Ions (Cations)							
1+		2+		3+		4+	
Ammonium	NH ₄ ⁺	Barium	Ba ²⁺	Aluminum	Al ³⁺	Tin(IV)	Sn ⁴⁺
Cesium	Cs ⁺	Calcium	Ca ²⁺	Antimony(III)	Sb ³⁺	(Stannic)	
Copper(I)	Cu ⁺	Chromium(II)	Cr ²⁺	Arsenic(III)	As ³⁺	Lead(IV)	Pb ⁴⁺
(Cuprous)		Cobalt(II)	Co ²⁺	Bismuth(III)	Bi ³⁺	(Plumbic)	
Francium	Fr ⁺	Copper(II)	Cu ²⁺	Boron	B ³⁺	Manganese(IV)	Mn ⁴⁺
Gold	Au ⁺	(Cupric)		Chromium(III)	Cr ³⁺		
Hydrogen	H ⁺	Iron(II)	Fe ²⁺	Cobalt(III)	Co ³⁺		
Lithium	Li ⁺	(Ferrous)		Iron(III)	Fe ³⁺		
Potassium	K ⁺	Lead(II)	Pb ²⁺	(Ferric)			
Rubidium	Rb ⁺	(Plumbous)		Titanium(III)	Ti ³⁺		
Silver	Ag ⁺	Magnesium	Mg ²⁺				
Sodium	Na ⁺	Manganese(II)	Mn ²⁺				
		Mercury(I)	Hg ₂ ²⁺				
		(Mercurous)					
		Mercury(II)	Hg ²⁺				
		(Mercuric)					
		Nickel(II)	Ni ²⁺				
		Tin(II)	Sn ²⁺				
		(Stannous)					
		Zinc	Zn ²⁺				

Negative Ions (Anions)					
1-		2-		3-	
Acetate	C ₂ H ₃ O ₂ ⁻	Carbonate	CO ₃ ²⁻	Arsenate	AsO ₄ ³⁻
Bromate	BrO ₃ ⁻	Chromate	CrO ₄ ²⁻	Borate	BO ₃ ³⁻
Bromide	Br ⁻	Dichromate	Cr ₂ O ₇ ²⁻	Nitride	N ³⁻
Chlorate	ClO ₃ ⁻	Hydrogen Phosphate	HPO ₄ ²⁻	Phosphate	PO ₄ ³⁻
Chloride	Cl ⁻	(Biphosphate)		Phosphite	PO ₃ ³⁻
Chlorite	ClO ₂ ⁻	Oxalate	C ₂ O ₄ ²⁻	Phosphide	P ³⁻
Cyanide	CN ⁻	Oxide	O ²⁻		
Dihydrogen phosphate	H ₂ PO ₄ ⁻	Peroxide	O ₂ ²⁻		
Fluoride	F ⁻	Sulfate	SO ₄ ²⁻		
Hydride	H ⁻	Sulfide	S ²⁻		
Hydrogen carbonate	HCO ₃ ⁻	Sulfite	SO ₃ ²⁻		
(Bicarbonate)		Silicate	SiO ₃ ²⁻		
Hydrogen oxalate	HC ₂ O ₄ ⁻	Thiosulfate	S ₂ O ₃ ²⁻		
Hydrogen sulfate	HSO ₄ ⁻				
(Bisulfate)					
Hydrogen sulfide	HS ⁻				
Hydrogen sulfite	HSO ₃ ⁻				
(Bisulfite)					
Hydroxide	OH ⁻				
Hypochlorite	ClO ⁻				
Iodide	I ⁻				
Nitrate	NO ₃ ⁻				
Nitrite	NO ₂ ⁻				
Perchlorate	ClO ₄ ⁻				
Permanganate	MnO ₄ ⁻				
Thiocyanate	SCN ⁻				